

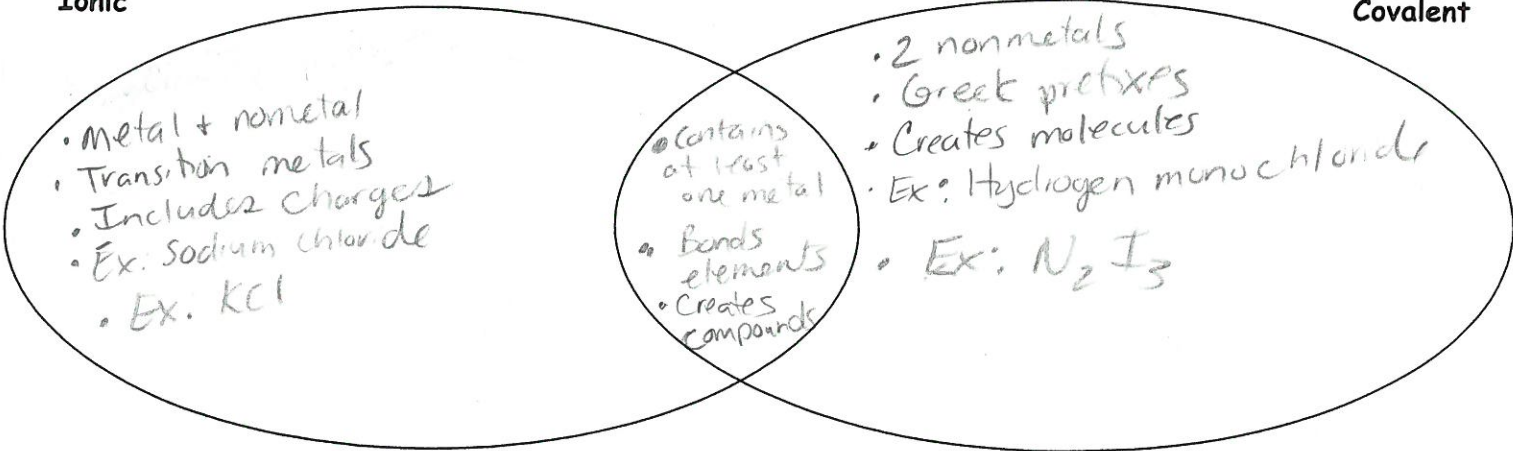
Key

CHEMICAL BONDING & CHEMICAL REACTIONS REVIEW

I. Chemical Bonds- Write the items shown on the screen in the appropriate place.

Ionic

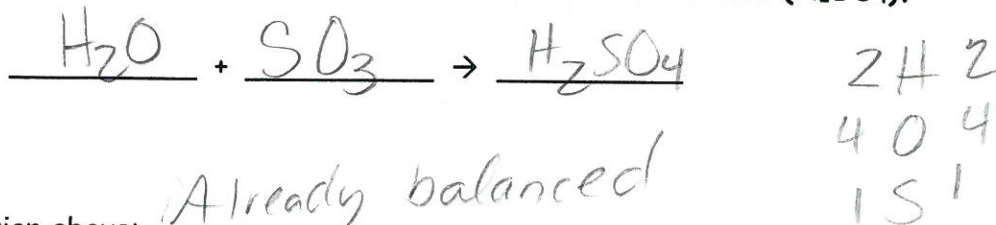
Covalent



II. Translate Words to Equations

1 a. Write the following reaction in word format:

Water and sulfur trioxide react to form sulfuric acid (H₂SO₄).



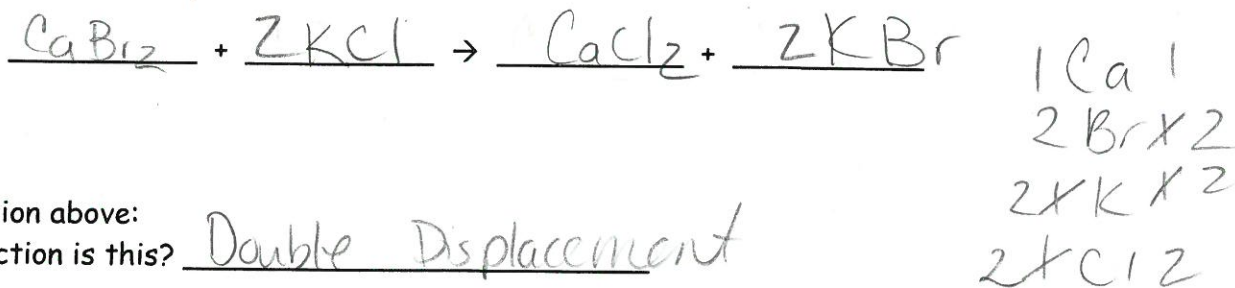
Already balanced

b. Balance the equation above:

c. What type of reaction is this? Synthesis

2 a. Write the following reaction in formula format:

Calcium bromide and Potassium chloride react to form Calcium chloride and Potassium bromide.



Double Displacement

b. Balance the equation above:

c. What type of reaction is this? Double Displacement

III. Reaction Types

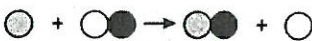
3. Type: Decomposition



4. Type: Synthesis



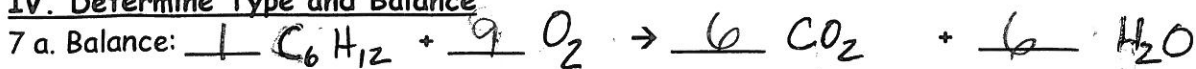
5. Type: single displacement



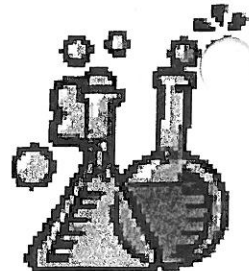
Type: Double displacement



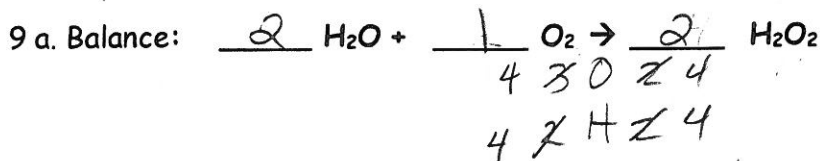
IV. Determine Type and Balance



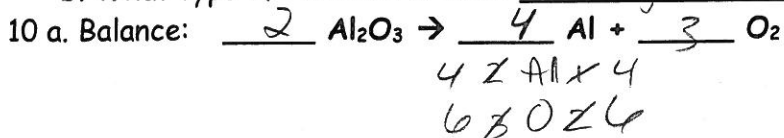
b. What type of reaction is this? Combustion



b. What type of reaction is this? Single displacement



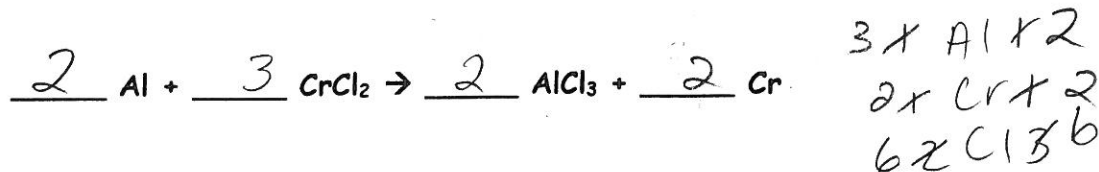
b. What type of reaction is this? Synthesis



b. What type of reaction is this? Decomposition

VI. Law of Conservation of Mass

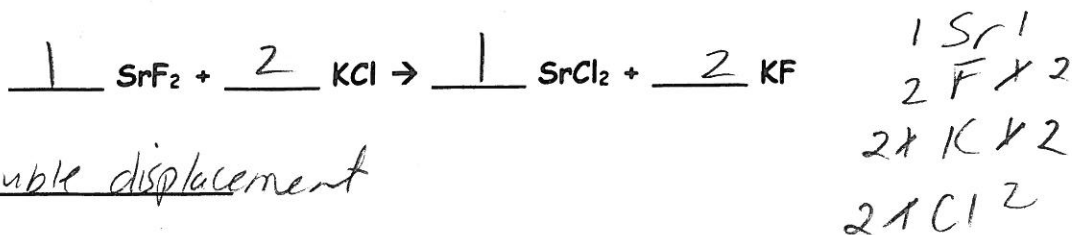
11 a. During your science lab, Aluminum was added to Chromium (II) chloride. Based on the Law of Conservation of Matter, if 15 grams of Aluminum are added to 25 grams of Chromium (II) chloride to produce 8 grams of Aluminum chloride. How much Chromium would also be produced? 32 grams



b. Balance the equation

c. Type of reaction: Single displacement

12 a. During a second science experiment, Strontium fluoride was added to Potassium chloride. Based on the Law of Conservation of Matter, if 4.5 grams of Strontium fluoride are added to 5.5 grams of Potassium chloride to produce 6 grams of Strontium chloride. How much Potassium fluoride would also be produced? 4 grams



b. Balance the equation

c. Type of reaction: double displacement