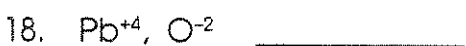
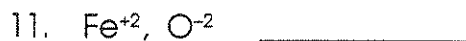


WRITING BINARY FORMULAS

Name _____

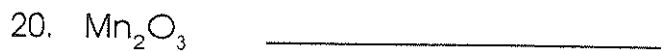
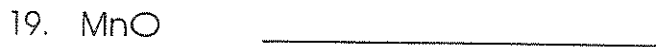
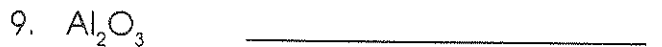
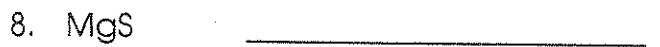
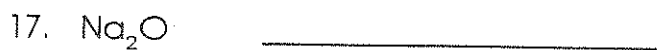
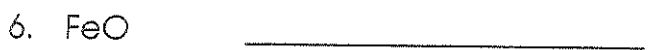
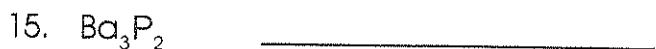
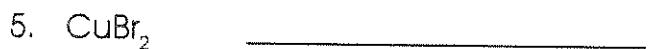
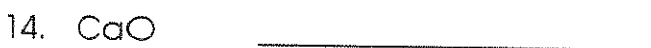
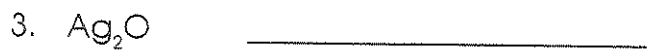
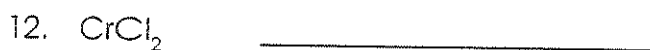
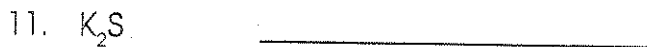
Write the formulas for the compounds formed from the following ions.



NAMING BINARY COMPOUNDS (IONIC)

Name _____

Name the following ionic compounds using Roman numerals where necessary.



FORMULAS WITH POLYATOMIC IONS

Name _____

Matching the horizontal and vertical axes, write the formulas of the compounds with the following combination of ions. The first one is done for you.

	OH^-	NO_3^-	CO_3^{2-}	SO_4^{2-}	PO_4^{3-}
H^+	HOH (H_2O)	HNO_3	H_2CO_3	H_2SO_4	H_3PO_4
Na^+					
Mg^{+2}					
NH_4^+					
Ca^{+2}					
K^+					
Al^{+3}					
Pb^{+4}					

NAMING OF NON-BINARY COMPOUNDS

Name _____

An ionic compound that contains more than two elements must contain a polyatomic ion. Name the following compounds.

1. NaNO_3 _____
2. Ca(OH)_2 _____
3. K_2CO_3 _____
4. NH_4Cl _____
5. MgSO_4 _____
6. AlPO_4 _____
7. $(\text{NH}_4)_2\text{SO}_4$ _____
8. Na_3PO_4 _____
9. CuSO_4 _____
10. NH_4OH _____
11. Li_2SO_3 _____
12. $\text{Mg(NO}_3)_2$ _____
13. Al(OH)_3 _____
14. $(\text{NH}_4)_3\text{PO}_4$ _____
15. KOH _____
16. $\text{Ca(NO}_3)_2$ _____
17. K_2SO_4 _____
18. Pb(OH)_2 _____
19. Na_2O_2 _____
20. CuCO_3 _____